# UNIVERSITI TUN HUSSEIN ONN MALAYSIA

## FINAL EXAMINATION
## SEMESTER I
## SESSION 2012/2013

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|---|---|---|
| COURSE NAME | : | INTRODUCTION TO CHEMICAL ENGINEERING TECHNOLOGY |
| COURSE CODE | : | BNQ 10102 |
| PROGRAMME | : | BNN |
| TEST DATE | : | JANUARY 2013 |
| DURATION | : | 2 HOURS |
| INSTRUCTION | : | ANSWER **THREE (3)** QUESTIONS ONLY |

THIS QUESTION PAPER CONSISTS OF **FOUR (4)** PAGES

**Q1** (a) Define the term *dimension* and give **three (3)** examples of base dimensions.

(5 marks)

(b) Describe the derived dimension of *volume* and of *density* in terms of the base dimensions.

(2 marks)

(c) Demonstrate which of these equations are dimensionally homogenous?

(i) $x(\text{m}) = x_o(\text{m}) + 0.304(\text{m/ft})v(\text{ft/s})t(\text{s}) + 0.5a(\text{m/s}^2)[t(\text{s})]^2$

(ii) $P(\text{kg/ms}^2) = 10135(\text{Pa/atm})\ 1(\text{kg/ms}^2/\text{Pa})P_o(\text{atm}) + p(\text{kg/m}^3)v(\text{m/s})$

(5 marks)

(d) Yeast for home bakers is sold in ¼ oz. packages. If one yeast cell weighs about 6x10-5 μg, calculate how many yeast cells are in a package?

(4 marks)

(e) Nitrogen gas costs 25 cents per 100 standard cubic feet. Liquid nitrogen costs 28 cents per liter. The specific gravity of liquid nitrogen is 0.808. Compare the costs of gas and liquid nitrogen on dollars per kilogram basis. Explain why is one so much cheaper than the other?

(9 marks)

**Q2** (a) Calculate the molar density of a gas at:

(i) Ideal conditions: 0°C and 1 atm

(ii) 100°C and 3.50 atm

(5 marks)

(b) 25 gmol glucose ($C_6H_{12}O_6$) sits in a beaker. Calculate how many grams of glucose, and how many moles of carbon are in the beaker? (Molar mass = 180 g/gmol glucose).

(6 marks)

(c) In 100 ml cup of water, demonstrate how many grams of H2O are there in the cup? How many moles?

(4 marks)

(d) The human body contains 63% H, 25.5% O, 9.45% C, 1.35% N, 0.31% Ca, and 0.22% of P, plus several elements present in trace.

(i) Are these in mass or mole percents?
(ii) About how many grams and moles of each element do you carry around?

(10 marks)

**Q3** (a) Briefly explain what is meant by these processes. Sketch a diagram for each process.

(i) Batch process
(ii) Continuous flow process
(iii) Semi-batch process

(12 marks)

(b) Your job is to design a mixer to produce **300 kg/day** of battery acid. The mixer will operate continuously and at steady state. The battery acid product must contain 18.6 wt% $H_2SO_4$ in water. Raw materials available include concentrated sulfuric acid solution at 77 wt% $H_2SO_4$ in water, and pure water. What is the flow rate of each raw material into the mixer? Illustrate your answer in a diagram

(13 marks)

**Q4** (a) Illustrate **three (3)** routes of entry of chemicals into human body.

(9 marks)

(b) Sketch **three (3)** symbols indicating different classes of hazardous substances.

(6 marks)

(c) Discuss **five (5)** effects of prolong chemical exposure to human.

(10 marks)

**- END OF QUESTION –**

## APPENDIX: PERIODIC TABLE

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|---|---|---|---|---|---|---|---|---|---|---|---|---|---|---|---|---|---|
| hydrogen 1 H 1.0079 | | | | | | | | | | | | | | | | | helium 2 He 4.0026 |
| lithium 3 Li 6.941 | beryllium 4 Be 9.0122 | | | | | | | | | | | boron 5 B 10.811 | carbon 6 C 12.011 | nitrogen 7 N 14.007 | oxygen 8 O 15.999 | fluorine 9 F 18.998 | neon 10 Ne 20.180 |
| sodium 11 Na 22.990 | magnesium 12 Mg 24.305 | | | | | | | | | | | aluminium 13 Al 26.982 | silicon 14 Si 28.086 | phosphorus 15 P 30.974 | sulfur 16 S 32.065 | chlorine 17 Cl 35.453 | argon 18 Ar 39.948 |
| potassium 19 K 39.098 | calcium 20 Ca 40.078 | | scandium 21 Sc 44.956 | titanium 22 Ti 47.867 | vanadium 23 V 50.942 | chromium 24 Cr 51.996 | manganese 25 Mn 54.938 | iron 26 Fe 55.845 | cobalt 27 Co 58.933 | nickel 28 Ni 58.693 | copper 29 Cu 63.546 | zinc 30 Zn 65.39 | gallium 31 Ga 69.723 | germanium 32 Ge 72.61 | arsenic 33 As 74.922 | selenium 34 Se 78.96 | bromine 35 Br 79.904 | krypton 36 Kr 83.80 |
| rubidium 37 Rb 85.468 | strontium 38 Sr 87.62 | | yttrium 39 Y 88.906 | zirconium 40 Zr 91.224 | niobium 41 Nb 92.906 | molybdenum 42 Mo 95.94 | technetium 43 Tc [98] | ruthenium 44 Ru 101.07 | rhodium 45 Rh 102.91 | palladium 46 Pd 106.42 | silver 47 Ag 107.87 | cadmium 48 Cd 112.41 | indium 49 In 114.82 | tin 50 Sn 118.71 | antimony 51 Sb 121.76 | tellurium 52 Te 127.60 | iodine 53 I 126.90 | xenon 54 Xe 131.29 |
| caesium 55 Cs 132.91 | barium 56 Ba 137.33 | 57-70 * | lutetium 71 Lu 174.97 | hafnium 72 Hf 178.49 | tantalum 73 Ta 180.95 | tungsten 74 W 183.84 | rhenium 75 Re 186.21 | osmium 76 Os 190.23 | iridium 77 Ir 192.22 | platinum 78 Pt 195.08 | gold 79 Au 196.97 | mercury 80 Hg 200.59 | thallium 81 Tl 204.38 | lead 82 Pb 207.2 | bismuth 83 Bi 208.98 | polonium 84 Po [209] | astatine 85 At [210] | radon 86 Rn [222] |
| francium 87 Fr [223] | radium 88 Ra [226] | 89-102 ** | lawrencium 103 Lr [262] | rutherfordium 104 Rf [261] | dubnium 105 Db [262] | seaborgium 106 Sg [266] | bohrium 107 Bh [264] | hassium 108 Hs [269] | meitnerium 109 Mt [268] | ununnilium 110 Uun [271] | ununnium 111 Uuu [272] | ununbium 112 Uub [277] | | ununquadium 114 Uuq [289] | | | | |

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| *Lanthanide series | lanthanum 57 La 138.91 | cerium 58 Ce 140.12 | praseodymium 59 Pr 140.91 | neodymium 60 Nd 144.24 | promethium 61 Pm [145] | samarium 62 Sm 150.36 | europium 63 Eu 151.96 | gadolinium 64 Gd 157.25 | terbium 65 Tb 158.93 | dysprosium 66 Dy 162.50 | holmium 67 Ho 164.93 | erbium 68 Er 167.26 | thulium 69 Tm 168.93 | ytterbium 70 Yb 173.04 |
| **Actinide series | actinium 89 Ac [227] | thorium 90 Th 232.04 | protactinium 91 Pa 231.04 | uranium 92 U 238.03 | neptunium 93 Np [237] | plutonium 94 Pu [244] | americium 95 Am [243] | curium 96 Cm [247] | berkelium 97 Bk [247] | californium 98 Cf [251] | einsteinium 99 Es [252] | fermium 100 Fm [257] | mendelevium 101 Md [258] | nobelium 102 No [259] |